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(54) Title: METHOD FOR PRODUCING CATALYST STRUCTURES

(57) Abstract: A process for making microporous structures that can be used as a catalyst support. The microporous structures have high porosity and high thermal stability, combined with good mechanical strength and relatively high surface area. The process is useful for making titanium dioxide for catalyst structures for use for fuel cells, sensors electrochemical cells and the like.



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METHOD FOR PRODUCING CATALYST STRUCTURES

[001] This application claims priority from U.S. Serial No. 60/230,211, the entire contents of which are incorporated herein by reference.

[002] The present invention relates to a process for making microporous structures for use as a catalyst support. In particular, the present invention relates to a process for making titanium dioxide suitable as a catalyst structure for fuel cells, sensors, electrochemical cells, and the like. The present invention also relates to a catalyst structure formed from the process of the present invention.

BACKGROUND OF THE INVENTION

[003] Different materials and processes are known for the manufacture of catalyst supports. High porosity and good physical strength are general requirements for such products. High temperature stability of the structure is also required if the catalyst operates at elevated temperature.

[004] U.S. Pat. No. 5,036,037 teaches a method to produce metal oxide catalysts by pyrohydrolysis from solutions of chlorides, fluorides, or nitrates. The process makes particles with a mean size of 20-30 microns and with a high specific surface area. The processing temperature is at least 500° C. and generally higher than 700°C., to avoid the presence of the anion (chloride or fluoride) in the oxide product. The product can be used as such or can be further treated to give it the required physical or chemical properties. Although this process is suitable for the intended product, improvements to the process are desired.

[005] Novel processes for the manufacture of titanium dioxide from aqueous solutions have been disclosed in PCT Publications WO 01/00530, WO 01/00531, and WO 01/12555, the relevant portions of which are incorporated herein by reference. In general, these applications describe the processing of an aqueous solution of a titanium salt by evaporation to produce an intermediate. The evaporation is conducted at a temperature higher than the boiling point of the solution, but lower than the temperature where significant crystal growth of an oxide phase occurs. In some embodiments, the

evaporation may be conducted at a temperature higher than the boiling point of the solution but lower than the crystallization temperature of the intermediate.

[006] In the case of titanium solutions, the temperature generally ranges from 120° to 350° C, and preferably from 200° to 250° C. The process is preferably conducted by spraying, and can be accomplished in a spray dryer. The spray drying process produces thin-filmed spheres or parts of spheres, with a diameter of 1 to 100 μm , and a shell thickness of 0.03 to 5 μm .

[007] After calcination and milling of these spheres or parts of spheres, and depending on the conditions of evaporation, the choice of additives, and the conditions of calcination, ultra-fine nano-sized TiO_2 or, alternatively, pigment grade TiO_2 can be obtained.

[008] There has been no suggestion, however, that such a process can economically and commercially produce catalyst structures made of metal oxides from salt solutions of the metals. The present invention is therefore directed to a process to economically produce catalyst structures or catalyst supports.

[009] Accordingly, the present invention teaches a process to produce catalysts or catalyst structures with high porosity, high specific surface area, high mechanical strength, and excellent thermal stability. In contrast to the method disclosed in U.S. Pat. No. 5,036,037, the method of the present invention uses lower temperature equipment for the first step of the process and adjunction of chemical control additives. The method uses a combination of spraying, pressing, and crystallization, which allows optimal control of the physical characteristics of the product.

SUMMARY OF THE INVENTION

[010] The present invention provides a process for making catalysts or catalyst structures that comprises mixing an aqueous solution of a metal salt and a chemical control agent to form an intermediate solution. The solution is preferably free of any precipitate.

[011] The intermediate solution is then evaporated to form an intermediate product. The evaporation is conducted under conditions to achieve substantially total evaporation. In particular, the evaporation is conducted at a temperature higher than

the boiling point of the feed solution but lower than the temperature where significant crystal growth occurs. The evaporation may be conducted at a temperature higher than the boiling point of the solution but lower than the crystallization temperature of the intermediate. In a particularly preferred embodiment, the intermediate is an amorphous solid formed as a thin film and preferably is spherical or part of a sphere.

[012] The term “substantially total evaporation” or “substantially complete evaporation” refers to evaporation such that the solid intermediate contains less than 15% free water, preferably less than 10% free water, and more preferably less than 1% free water. The term “free water” is understood and means water that is not chemically bound and can be removed by heating at a temperature below 150° C. After substantially total evaporation or substantially complete evaporation, the intermediate product will have no visible moisture present.

[013] The intermediate product is then mixed with a binder to form a mixture that is then dried. The drying can be performed in any suitable manner but is preferably air dried. The dried mixture is then pressed into a desired shape. Suitable desired shapes include, but are not limited to disks, full cylinders or hollow cylinders in the size range of a few mm to 20 cm.

[014] The shaped product may then be heated to a temperature of 100° C to remove any remaining moisture or volatile compounds and prevent cracks during the heat treatment (crystallization) step.

[015] The heat treated product is then crystallized by raising the temperature to a temperature between 500°C to 1300°C for a period of time from 2 to 24 h and then cooled to room temperature. The cooled product is then washed by immersing it in water or dilute acid, heated to boiling and maintained at the boiling point for a period of time from 5 min to 2 h, to remove traces of any water-soluble phase that may still be present after the crystallization step.

[016] The product formed as a result of the process of the present invention includes a catalyst structure characterized by a porosity in the range of 30% to 70% and a thermal stability such that less than 5% dimensional change occurs upon holding the structure at a temperature of 1100° C in an oxidizing atmosphere for 8 hours. A

preferred product is a titanium dioxide catalyst structure consisting of needle-shaped particles that are strongly bound together while exhibiting a high porosity.

BRIEF DESCRIPTION OF THE DRAWINGS

[017] FIG. 1 is a general flow sheet of the process according to the present invention.

[018] FIG. 2 is a scanning electron micrograph of one example of a catalyst structure obtained according to the process of the present invention.

[019] FIG. 3 is a scanning electron micrograph of another example of a catalyst structure obtained according to the process of the present invention.

DESCRIPTION OF THE INVENTION

[020] Referring now to FIG 1, a generalized flow sheet of the process of the present invention is shown. With regard to FIG. 1, it has been found that some of the same processing steps and some of the equipment described in U.S. Patent Application Serial Nos. 09/500,207, 09/503,365 and 09/503,636, the entire contents of which are incorporated by reference, can advantageously be used to economically produce catalyst structures made of metal oxides, from salt solutions of these metals. In this regard, it is noted that these patent applications correspond to PCT publications WO 01/00531, WO 01/12555, and WO 01/00530, respectively. The contents of these publications are described above.

[021] In accordance with one embodiment of the process of the present invention, a metal salt solution 2 is mixed with one or more chemical control agents 4 in mixer 10. The metal salt solutions are generally sulfates, chlorides, oxychlorides, nitrates, or mixtures thereof. In this regard, the metal forming the salt can be selected from the group consisting of Ti, Zr, Hf, V, Nb, Ta, Cr, Mo, W, Mn, Al, Sn, Sb, Pb, Bi, and mixtures thereof. The metal salt solution may include the soluble salts of Y, Ca, Mg, and mixtures thereof. For example, when the metal forming the salt is Ti or Zr, the metal salt solution may desirably include the soluble salts of Y, Ca, Mg. and mixtures thereof. The solution is preferably free of any precipitate or suspension.

[022] The process is particularly suited to aqueous solutions of titanium and zirconium, more particularly solutions of titanium oxychloride and zirconium oxychloride.

The chemical control agents can be, but are not limited to chloride salts of sodium, potassium, lithium, aluminum, tin and zinc. Carbonate, fluoride, sulfate, silicate, borate and other suitable salts of the aforementioned elements may also be used. Additionally, phosphoric acid and phosphate salts of the aforementioned elements may be used. Accordingly, the chemical control agent is selected from the group consisting of chloride salts, carbonate salts, fluoride salts, sulfate salts, silicate salts, borate salts and phosphate salts of sodium, potassium, lithium, aluminum, tin, and zinc and phosphoric acid. In a preferred process, titanium oxychloride solution is used as feed and sodium phosphate is used as a chemical control agent.

[023] The mixed solutions are subjected to substantially total evaporation 20 by contact with a hot surface or by spraying in a stream of hot gas to form an intermediate product. The substantially total evaporation is conducted above the boiling point of the solution and below the temperature where there is significant crystal growth. The intermediate product is an amorphous hydrous oxide. The term "substantially total evaporation" means that the resulting intermediate product contains less than 15% free water, preferably less than 10% free water, and more preferably less than 1% free water. The term "free water" is understood and means water that is not chemically bound and can be removed by heating at a temperature below 150°C. After substantially total evaporation, the intermediate product will have no visible moisture present. Water and volatile products of the acid involved are vaporized and may be recovered by any known process.

[024] Preferably, evaporation by the spraying process is accomplished in a spray dryer. The use of a spray dryer allows the resulting physical and chemical characteristics of the solid product to be controlled within a fairly narrow range through control of the operating parameters, including temperature, flow rate, concentration of the metal, and the type and amount of chemical control agents. In general, the temperature in the spraying chamber is in the range of 100° C. to 400° C. and the concentration in the feed solution is in the range of 10 to 200 g/l of metal. Preferably, the temperature range is from 150° C. to 250° C and the concentration range of the metal is from 50 to 150 g/l.

[025] The intermediate product resulting from spraying in a spray dryer will be composed of thin-filmed spheres or parts of spheres. The dimensions of the spheres may vary over a wide range, from 1 μm to 100 μm in diameter, and the shell thickness in the range from 30 nm to 5000 nm, preferably from 30 nm to 1000 nm. This intermediate product may be called amorphous.

[026] This intermediate product of the spraying process is further mixed with a binder 6 in the second mixing step 30 and optionally with surface treatment additives 8. The binder may include material effective to provide surface treatment of the intermediate product such as surface deposition and surface modification. It may involve complex microstructure control that may be used to enhance desired properties (i.e. thermal catalytic properties, conductivity of the surface layer or morphology of the microstructure). The binders have a lubricating action during the pressing process and are necessary to provide a well-finished surface. Binders must have good burnout properties during the heat treatment process.

[027] The binder may be an organic or an inorganic binder. Inorganic binders may include but are not limited to metal salt solutions, salts, colloidal metals and complex mixtures thereof. Organic binders may be natural or synthetic. Natural organic binders may include but are not limited to starch and its derivatives, cellulose ethers such as caboxymethyl cellulose, hydroxyethyl cellulose, methyl hydroxypropyl cellulose and soybean protein. Synthetic organic binders may include but are not limited to polymers or copolymers of styrene, butadiene, acryl ester, vinyl acetate, and acrylonitrile. Polyvinyl alcohol, ammonium polyacrylate and sodium polyacrylate are examples of suitable organic binders.

[028] The optional surface treatment additives are generally similar to the chemical control agents. Accordingly, the optional surface treatment additives include, but are not limited to, chloride salts of sodium, potassium, lithium, aluminum, tin and zinc. Carbonate, fluoride, sulfate, silicate, borate and other suitable salts of the aforementioned elements may also be used. Additionally, phosphoric acid and phosphate salts of the aforementioned elements may be used. Accordingly, the chemical control agent is selected from the group consisting of chloride salts, carbonate

salts, fluoride salts, sulfate salts, silicate salts, borate salts and phosphate salts of sodium, potassium, lithium, aluminum, tin, and zinc and phosphoric acid.

[029] Thereafter, the product is dried in a drying step 40. Drying may be performed in any suitable manner. The material to be dried may be placed on shelves in a drying oven, or it may be passed in continuous motion through a belt oven or a pusher oven. Drying may also be accomplished in a rotary kiln. Heating may be provided for example by an electric heating resistance, by a flow of hot air or by a flow of hot combustion gases.

[030] The dried material is pressed into the desired shape in a pressing step 50. The pressure applied may vary over a wide range, but is preferably in the range from 2000 to 20,000 psi. A hydrostatic press may be used to press disks or cylinders. Special attachments may be used to make pressed material of different shapes.

[031] The shaped product may then be heated to a temperature of 100° C to remove any remaining moisture or volatile compounds and prevent cracks during the heat treatment (crystallization) step. The pressed product is then subjected to a heat treatment step 60 where the pressed product is further dried and slowly heated. In a preferred embodiment, the heat treatment is conducted over a period of at least 2 h to a temperature where crystallization and crystal growth occurs.

[032] This crystallization temperature varies with the nature of the material serving as a catalyst structure, as well as with the nature of the addition. The temperature and conditions of conversion are dependent upon the nature and amount of the additives. The temperature is generally in the range of 500°C to 1300°C for a period of time from 2 to 24 h and then cooled to room temperature. The temperature is generally in the range of 800° C. to 1200° C for TiO₂ and in the range from 1000°C to 1300°C for ZrO₂. The total time of a heat treatment cycle varies from 2 hours to 24 hours.

[033] The equipment used for heat treatment may be any kind of furnace with good temperature control. Small productions may be made in a muffle furnace. For larger productions, a continuous belt furnace with zones at different temperatures is preferred.

[034] After this heat treatment step, the product may be subjected to one or more washing steps 70, which may include washing with dilute acid and/or washing with water to remove soluble traces of the chemical control agent that was introduced before the spray-drying step or to remove the optional surface treatment additive introduced after evaporation but prior to the heat treatment step.

[035] The washing steps may include a step where the heat treatment product and the washing solution are heated to boiling and maintained at the boiling point for a period of time from 5 min to 2 h, to remove traces of any water-soluble phase that may not have reacted during the heat treatment step.

[036] The macrostructure of the product may be further modified by additional surface treatments or depositions 80. A different metal oxide may be deposited by chemical vapor deposition or from liquid phase solutions of the metal oxide precursor. This coated material may then be thermally treated to mineralize the surface coating. The coating and curing process may be repeated several times until the desired coating thickness is developed.

[037] Without being bound by any theory, it is believed that the spraying process under the conditions of the present process yields spheres composed of a thin film of an amorphous solid that can be converted to a crystal structure with the required properties. Particularly, the crushed and compacted spherical shells create, after crystallization, a structure that is both very porous and mechanically robust.

[038] The process of the present invention produces catalyst or crystalline structures with a porosity in the range from 30% to 70%. In addition, the structures formed from the process of the present invention are characterized by having a thermal stability such that less than 5% dimensional change occurs upon holding the structure at 1100° C. in an oxidizing atmosphere for 8 hours. The specific surface area may vary over a wide range, but is generally in the range of 1 to 10 m²/g. The mechanical crushing strength of the catalyst structure produced is at least 1 MPa (10 bar or approximately 145 psi). The size of the individual particles forming the porous structure is generally in the micron range, from 0.1 micron to 50 micron in length. The particles may have any shape, but are often needles with a width to length ratio from 1:1 to 1:20.

[039] The size of these structures is not limited by the mentioned techniques. Further, the surface may be modified to include metal or metal oxides or may be coated with other metal oxides and solid solutions. The process of the present invention also allows the creation of multilayered crystalline materials fused into a porous macrostructure that may be easily swept by reactive gases or solutions.

[040] The following examples illustrate the present invention.

EXAMPLE I

[041] A solution of titanium chloride containing 40 g/l Ti, 140 g/l Cl and 0.2% $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$ was injected and subjected to substantially total evaporation in a spray dryer. The spray dryer consists of a reaction chamber followed by bag filters and a hydrochloric acid absorption system. The solution was injected at a rate of 3 liters/min through an atomizing disk. Gases from the combustion of natural gas, diluted with air to 580° C were also injected around the disk. The outlet temperature of the spray dryer was 250° C and the total gas flow about 800 scfm. Reactor off-gases were sent to a bag filter to collect the TiO_2 product.

[042] Ten grams of polyvinyl alcohol of molecular weight between 11,000 and 31,000 was dissolved in 100 ml of water. The mixture was further diluted into 200 ml of ethanol, and combined with 100 g of the titanium dioxide spray dryer product. The resulting paste was dried in a drying oven at 90° C, and then pressed into discs of 3.4 cm diameter under a pressure of 500 bar. The discs were further heat-treated according to the following cycle:

4 h at 110° C.

gradual temperature increase of 5° C/min up to 1150° C.

4 h at 1150° C.

the furnace was turned off and the product allowed to cool

The discs were washed by immersing them in water and brought to boiling for a period of 1 h.

[043] The resulting product showed a monolithic porous structure consisting of rutile crystals grown in place. This structure is clearly distinct from what would be obtained by sintering of a powdered crystalline material. XRD analysis showed only

pure rutile as the predominant solid phase and dimensional and density measurements indicated a void space fraction of 0.59.

[044] FIG. 2 is a scanning electron micrograph of the structure, enlarged 3000 times. It shows a structure of elongated rutile crystals with a cross-section of 1-2 μm and a length of 5 to 10 μm .

EXAMPLE II

[045] To a solution of titanium chloride containing 40 g/l Ti and 140 g/l Cl were added 0.67 g/l sodium silicate $\text{Na}_2\text{SiO}_3 \cdot 9\text{H}_2\text{O}$. The resultant solution was injected and subjected to substantially total evaporation in a spray dryer in the same conditions as those given in Example I. The titanium oxide product was recovered in a baghouse.

[046] Further treatment was identical to the treatment corresponding to Example I, except that the crystallization temperature was 920 °C instead of 1150°C.

[047] The resulting product showed a monolithic porous structure consisting of rutile crystals grown in place. XRD analysis showed only pure rutile as the solid phase. FIG. 3 is a scanning electron micrograph of the structure, enlarged 3000 times. It shows elongated rutile crystals with a length of 2 to 10 μm and a width to length ratio up to 1:10. The addition of sodium silicate generally produced more elongated crystals than the addition of sodium phosphate.

What is claimed:

1. A process for the manufacture of catalytic structures comprising:
 - a. mixing an aqueous solution of a metal salt and a chemical control agent to form an intermediate solution;
 - b. evaporating the intermediate solution to form an intermediate product wherein the evaporating is conducted in a controlled temperature process at a temperature higher than the boiling point of the solution but lower than the temperature where significant crystal growth occurs;
 - c. mixing the intermediate product with a binder to form a mixture;
 - d. drying the mixture to form a dried mixture;
 - e. pressing the dried mixture into a desired shape;
 - f. crystallizing by raising the temperature to a range between 500° C. to 1300° C. for a period of time from 2 to 24 h and thereafter by cooling to room temperature; and,
 - g. washing the product of step f.
2. The process of claim 1 wherein the metal forming the metal salt can be selected from the group consisting of Ti, Zr, Hf, V, Nb, Ta, Cr, Mo, W, Mn, Al, Sn, Sb, Pb, Bi, and mixtures thereof.
3. The process of claim 1 wherein the metal salt is selected from the group consisting of titanium oxychloride or zirconium oxychloride.
4. The process of claim 1 wherein the metal salt is titanium oxychloride and wherein during the crystallizing step the temperature is raised to a range between 800° C to 1200°C.
5. The process of claim 1 wherein the metal salt is zirconium oxychloride and wherein during the crystallizing step the temperature is raised to a range between 1000° C to 1300°C.

6. The process of claim 1 wherein the chemical control agent is selected from the group consisting of chloride salts, carbonate salts, fluoride salts, sulfate salts, silicate salts, borate salts and phosphate salts of sodium, potassium, lithium, aluminum, tin, and zinc and phosphoric acid.
7. The process of claim 1 wherein the evaporation step is conducted at a temperature between 100° C. and 400° C.
8. The process of claim 1 wherein the evaporating is conducted by spraying.
9. The process of claim 8 wherein the intermediate product comprises a plurality of hollow spheres and parts of spheres.
10. The process of claim 9 wherein the diameter of the spheres is between 1 μm and 100 μm .
11. The process of claim 9 wherein the thickness of the sphere is between 30 nm and 5000 nm.
12. The process of claim 1 wherein the binder is selected from the group consisting of inorganic binders, organic binders, and mixtures thereof.
13. The process of claim 1 wherein the washing is conducted by successively immersing the product in water, heating it to boiling, and keeping it at the boiling point for a period of time from 5 min to 2 h.
14. The process of claim 1 wherein a surface treatment additive is mixed with the intermediate product and binder.

15. A titanium dioxide catalyst structure characterized by a porosity in the range of 30% to 70% and a thermal stability such that less than 5% dimensional change occurs upon holding the structure at 1100° C. in an oxidizing atmosphere for 8 h.
16. The catalyst structure of claim 15 wherein the structure comprises needle-shaped particles that are strongly bound together.
17. A catalyst structure made according to the process of claim 1 wherein the structure has a porosity in the range of 30% to 70%.
18. The catalyst structure of claim 17 wherein the structure has a thermal stability such that less than 5% dimensional change occurs upon holding the structure at 1100° C. in an oxidizing atmosphere for 8 h.
19. The catalyst structure of claim 17 comprising a plurality of individual particles forming the structure wherein the particles have a size in a longitudinal direction from 0.1 to 50 micron.
20. The catalyst structure of claim 19 wherein the particles have a width to length ratio from 1:1 to 1:20.

FIG. 1

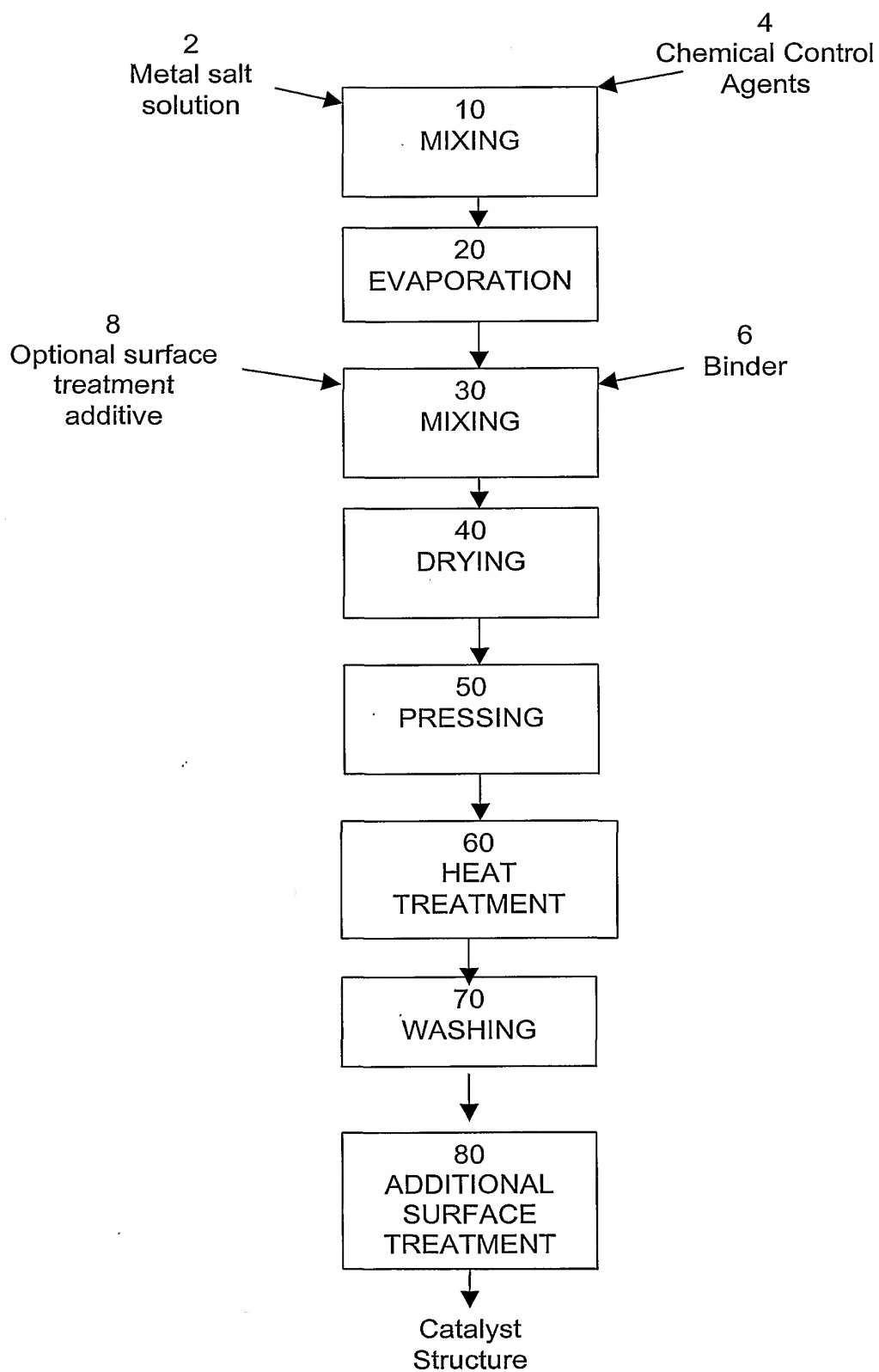


FIG. 2

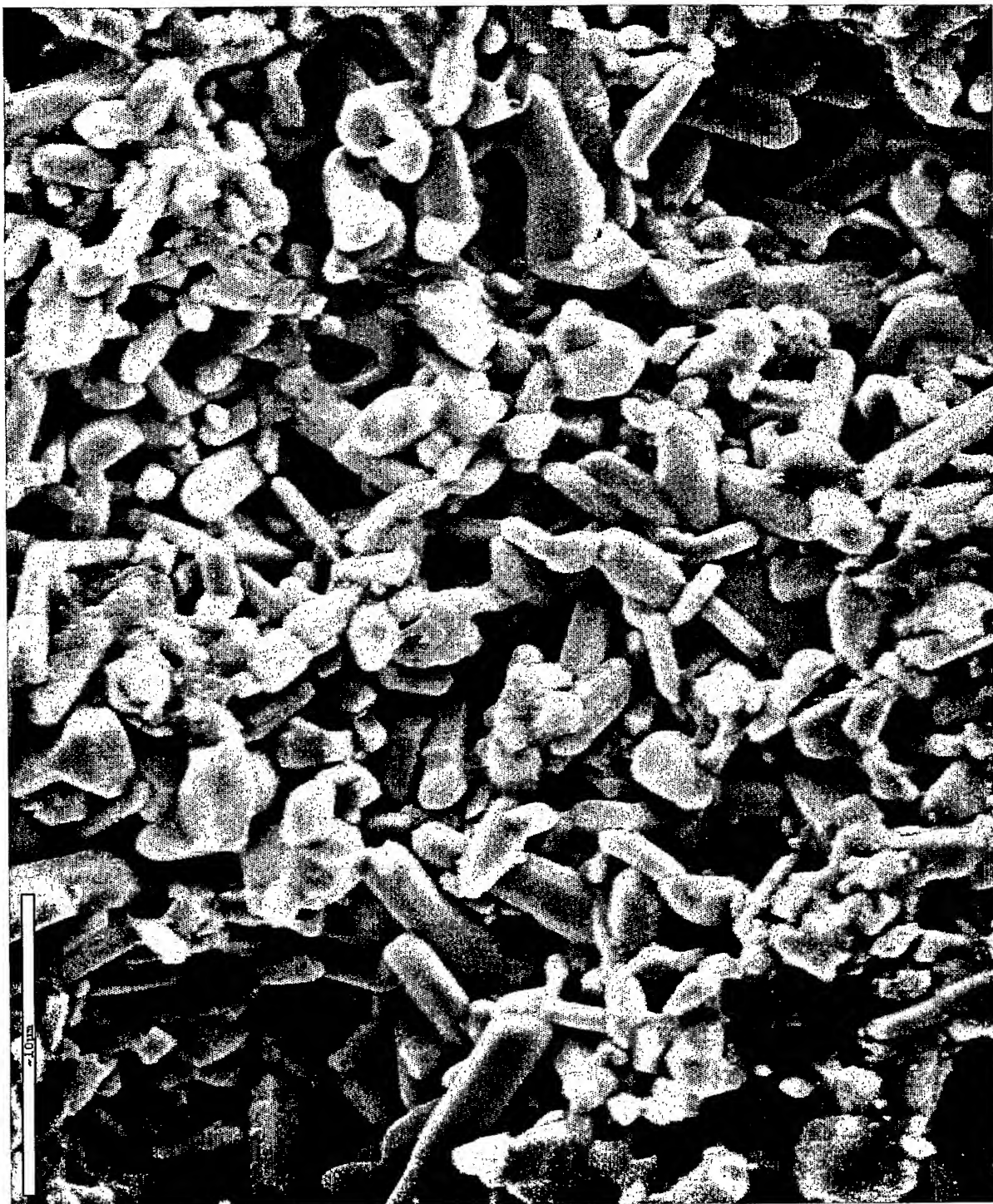


FIG. 3

